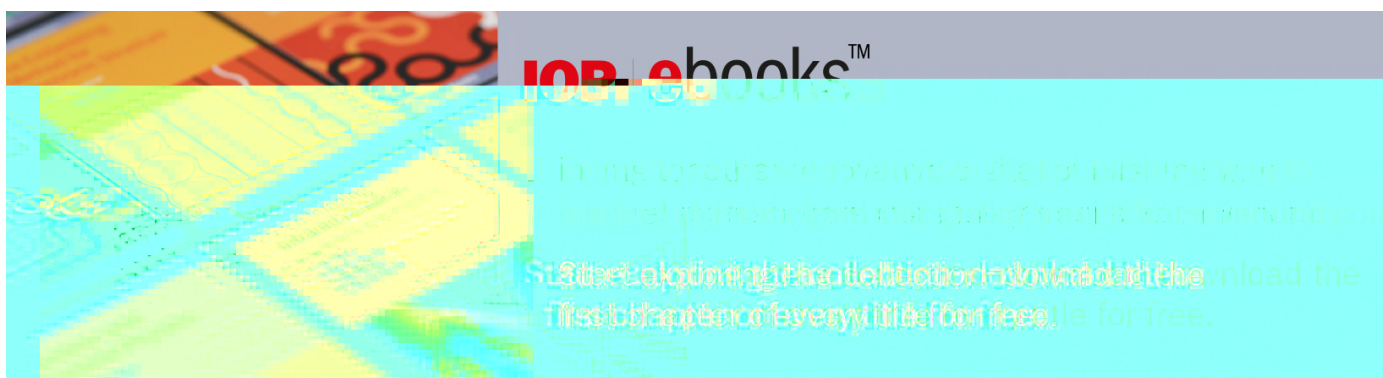


PAPER

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PAPER

Evolution of calcite surface reconstruction and interface adsorption of calcite-CO₂ with temperature

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Keywords: MD simulation, calcite, CO₂, interface adsorption

Abstract

Molecular dynamics (MD) simulation was used to study the evolution of calcite surface reconstruction and interface adsorption of calcite-CO₂ with temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature. At 673 K, the calcite surface reconstruction is significantly affected by the temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature.

1. Introduction

Calcite is a common mineral in the Earth's crust. It is a major component of many rocks and minerals. The study of calcite surface reconstruction and interface adsorption is of great importance for understanding the geochemical processes. In this paper, we use molecular dynamics (MD) simulation to study the evolution of calcite surface reconstruction and interface adsorption of calcite-CO₂ with temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature. At 673 K, the calcite surface reconstruction is significantly affected by the temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature. The results show that the calcite surface reconstruction is significantly affected by the temperature.

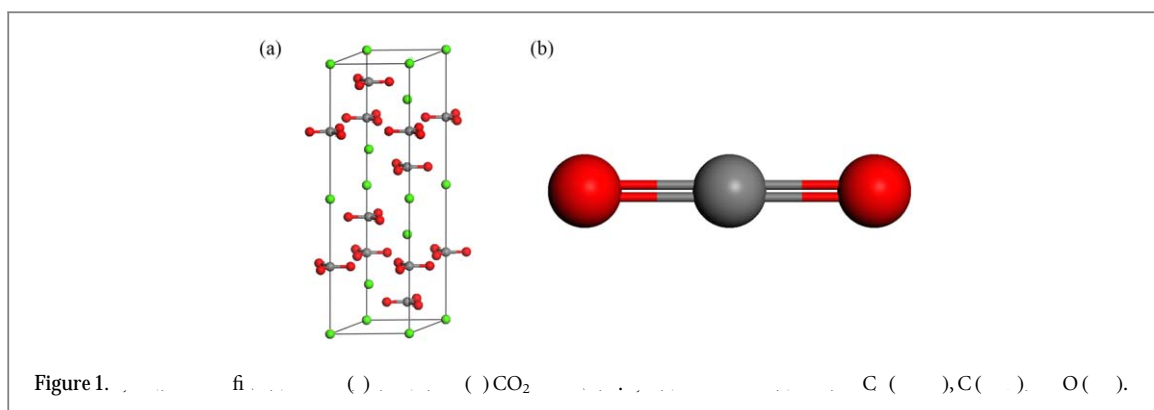


Table 1. χ^2 values for different matrix sizes (10³ /).

	E (×)						
	2 × 2	3 × 3	4 × 4	5 × 5	6 × 6	7 × 7	8 × 8
E	−1.464 ± 0.012	−1.447 ± 0.017	−1.451 ± 0.010	−1.445 ± 0.009	−1.445 ± 0.013	−1.447 ± 0.011	−1.445 ± 0.015



Figure 2. (a) Snapshot of a lipid bilayer (POPC) in a water box (5 × 5 × 3 nm) with a vacuum layer (4 nm thick) at the top. The lipid tails are shown in yellow and orange, and the headgroups are shown in red and blue. The water molecules are shown in red and white. The vacuum layer is shown in dark blue. The lipid bilayer is shown in the center of the water box. The lipid tails are shown in yellow and orange, and the headgroups are shown in red and blue. The water molecules are shown in red and white. The vacuum layer is shown in dark blue. The lipid bilayer is shown in the center of the water box.

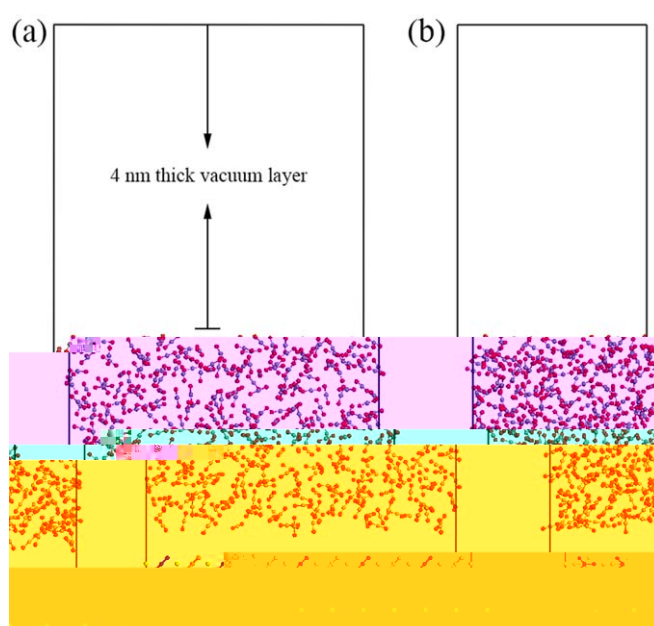


Figure 3. (a) Schematic diagram of a lipid bilayer (POPC) in a water box (5 × 5 × 3 nm) with a vacuum layer (4 nm thick) at the top. (b) Schematic diagram of a lipid bilayer (POPC) in a water box (5 × 5 × 3 nm) with a vacuum layer (4 nm thick) at the top. The lipid tails are shown in yellow and orange, and the headgroups are shown in red and blue. The water molecules are shown in red and white. The vacuum layer is shown in dark blue. The lipid bilayer is shown in the center of the water box.

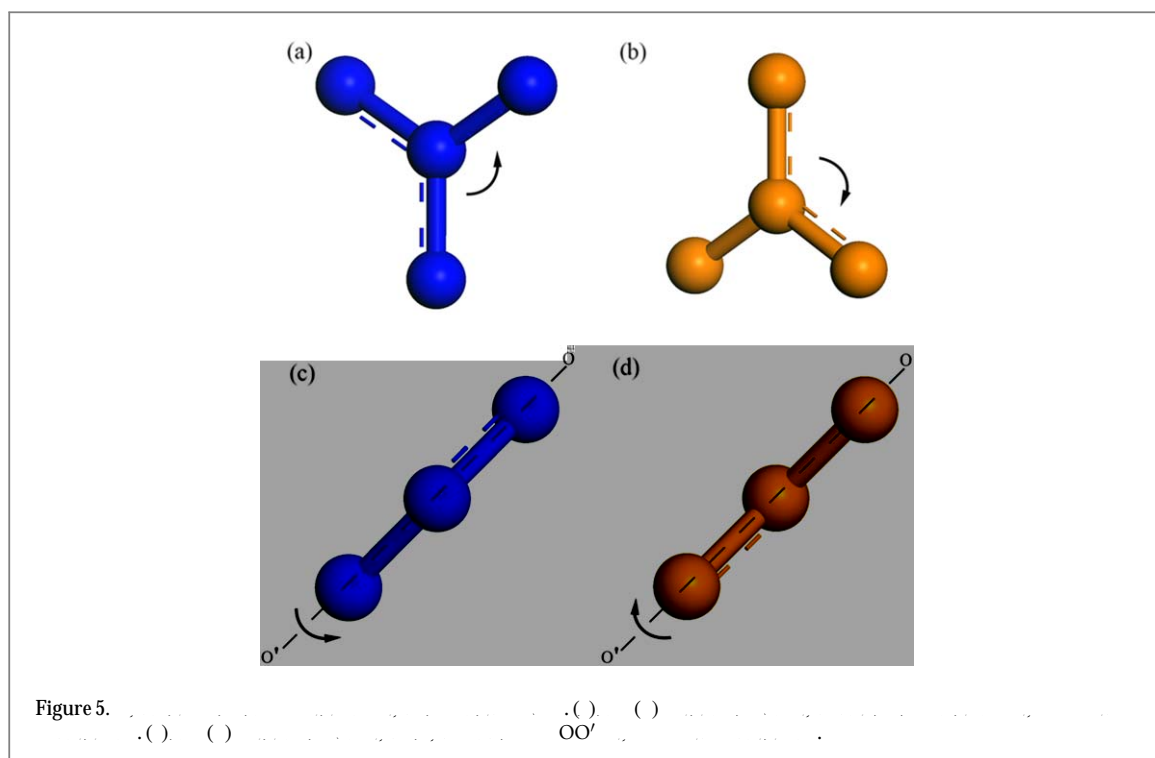


Table 2. E

(K)	(°)				O-C-O (°)		D (Å)		$\angle CO_3^{2-}$ (°)	D (Å)	Δ_{surface} (Å)
	R_C^B	R_C^Y	$R_{OO'}^B$	$R_{OO'}^Y$	\angle_{O-C-O}^B	\angle_{O-C-O}^Y	D_{Ca-O}^B	D_{Ca-O}^Y			
298	18.5	19.4	14.2	14.7	122.7	122.8	0.194	0.194	39.4		0.0669
373	19.3	19.1	15.1	14.3	123.0	122.9	0.194	0.195	39.3		0.0727
473	20.0	20.8	15.3	15.4	123.0	122.9	0.196	0.196	39.0		0.0721
573	22.6	21.1	16.6	16.1	123.3	123.0	0.197	0.196	38.9		0.0736
673	24.2	24.0	18.1	18.3	123.5	123.5	0.198	0.198	38.4		0.0715
773	24.3	24.7	18.6	18.7	123.6	123.5	0.198	0.198	38.3		0.0748
873	24.7	24.5	18.5	18.4	123.5	123.6	0.198	0.199	38.4		0.0771

OO' fl. 5.I $R_{OO'}^B$ $R_{OO'}^Y$

.B, \angle_{O-C-O}^B \angle_{O-C-O}^Y [16] O-C-O O

3° C-O C-O [1,17]. [45].

(298 K). I (104)

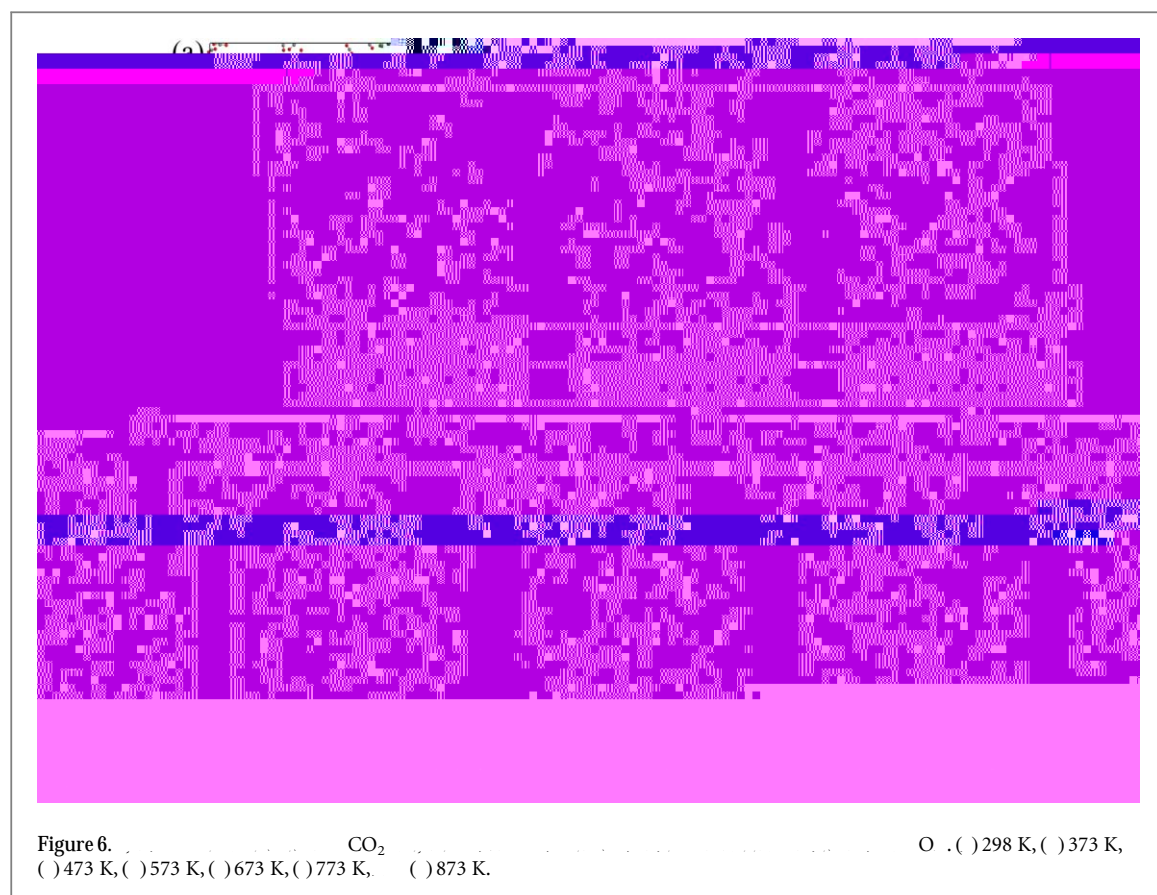
fl G *etal*[46]M fl (104)

.N, [47,48], [1].

C (D_{Ca-O} 0.231 [49]), C O.

C O. I, \angle_{O-C-O}^B \angle_{O-C-O}^Y [50]. C CO₃ [51].

C CO₃ [22].

Table 3.E
CO₂

(K)	$\angle \text{O-C-O}$ ($^\circ$)				$\angle \text{O-C-O}$ ($^\circ$)		D (\AA)		$\angle \text{CO}_3^{2-}$ ($^\circ$)	D (\AA)	Δ_{surface} (\AA)
	R_C^B	R_C^Y	$R_{OO'}^B$	$R_{OO'}^Y$	$\angle_{\text{O-C-O}}^B$	$\angle_{\text{O-C-O}}^Y$	$D_{\text{Ca-O}}^B$	$D_{\text{Ca-O}}^Y$			
298	18.5	19.4	14.4	14.7	122.6	122.7	0.193	0.194	39.6		0.0670
373	19.4	19.2	15.1	14.3	123.0	122.9	0.194	0.196	38.7		0.0722
473	20.1	20.9	15.4	15.4	123.1	123.0	0.196	0.196	39.0		0.0734
573	22.6	21.4	16.6	16.1	123.3	123.0	0.197	0.197	38.9		0.0740
673	24.1	23.8	18.0	17.9	123.5	123.4	0.198	0.197	38.1		0.0720
773	24.3	24.4	18.2	18.1	123.6	123.5	0.198	0.198	37.8		0.0748
873	24.5	24.6	18.3	18.4	123.5	123.6	0.197	0.198	37.9		0.0769

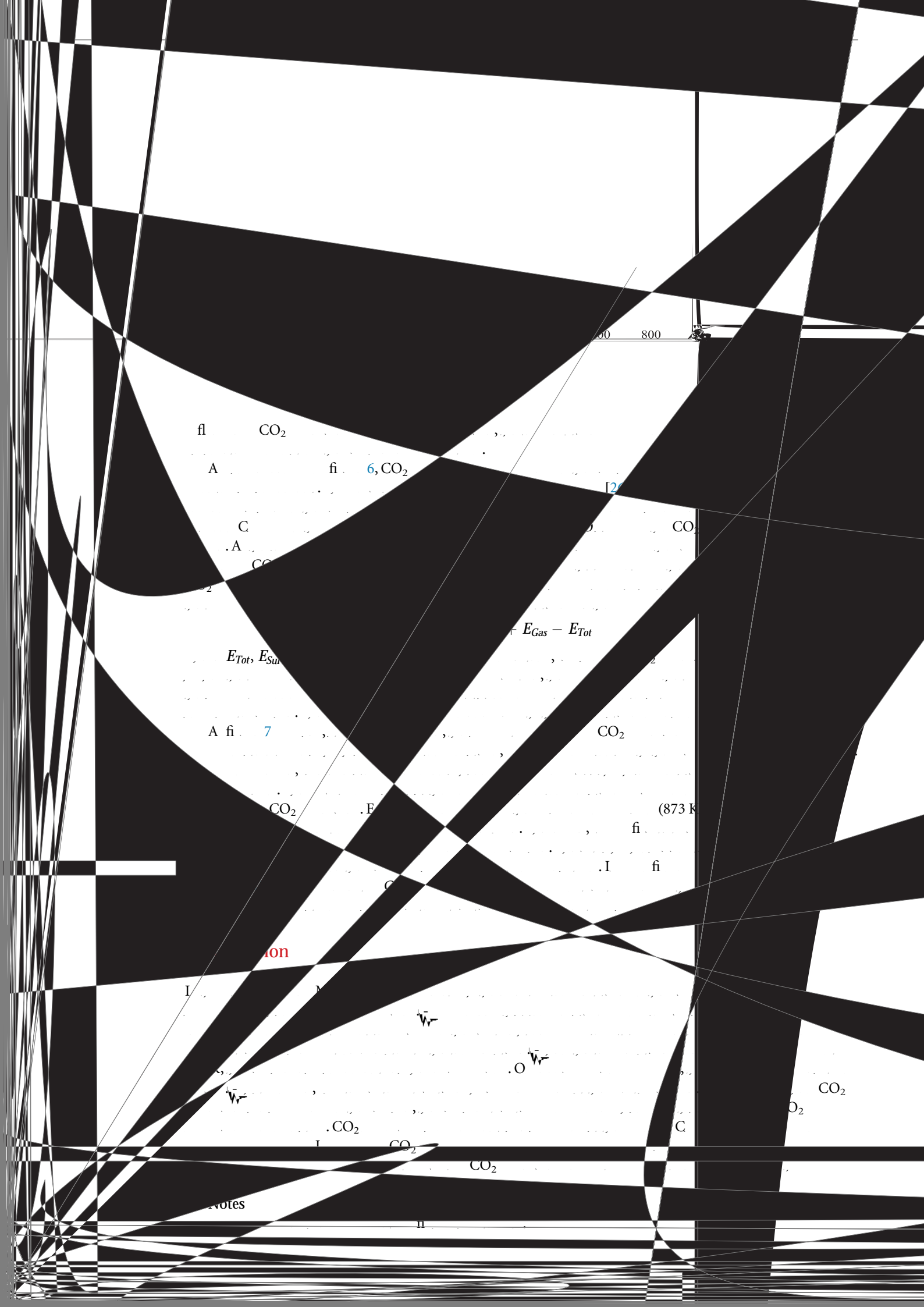
A ^{13}C NMR spectrum of the MD polymer, showing a broad peak at $\delta = 67.3$ ppm, assigned to the $\text{O}-\text{C}-\text{O}$ group, and a broad peak at $\delta = 161.1$ ppm, assigned to the $\text{C}=\text{O}$ group.

3.2. CO₂ adsorption behavior

$$I = \frac{1}{2} \int_{-L}^L \left[\frac{1}{2} \left(\frac{\partial \psi}{\partial x} \right)^2 + \frac{1}{2} \left(\frac{\partial \psi}{\partial y} \right)^2 + \frac{1}{2} \left(\frac{\partial \psi}{\partial z} \right)^2 + \frac{1}{2} \left(\frac{\partial \psi}{\partial t} \right)^2 \right] dx dy dz$$

Table 4. ΔE_{ads} and ΔE_{pure} of the adsorbent (10³ J / mol).

	Temperature (K)						
	298	373	473	573	673	773	873
E_{Pure}	−109 451 ± 75	−109 411 ± 81	−109 313 ± 84	−109 237 ± 78	−109 199 ± 85	−109 101 ± 92	−109 005 ± 95
E_{Ads}	−109 442 ± 72	−109 424 ± 77	−109 309 ± 80	−109 228 ± 79	−109 203 ± 88	−109 086 ± 97	−108 997 ± 94



200 800

fl CO₂

A fi 6, CO₂

[2]

C

.A

CO

2

$E_{Gas} - E_{Tot}$

E_{Tot}, E_{Sul}

A fi 7

CO₂

CO₂

.F

(873 K

fi

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M

\vec{w}

.O \vec{w}

\vec{w}

.CO₂

CO₂

CO₂

CO₂

O₂

C

Notes

ii

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